

A blue precipitate appears when two drops of dilute NaOH solution is added to a blue solution.

Name the ion present.

Give the formula for the precipitate.

No precipitate appears when two drops of dilute NaOH solution is added to a colourless solution. More NaOH is added and when heated the gas emitted turns red litmus paper blue.

Name the ion present.

Give its formula.

An orange/brown precipitate appears when two drops of dilute NaOH solution is added to an orange solution.

Name the ion present.

Give the formula for the precipitate.

A green precipitate appears when two drops of dilute NaOH solution is added to a pale green solution.

Name the ion present.

Give the formula for the precipitate.

A brown/grey precipitate appears when two drops of dilute NaOH solution is added to a colourless solution.

Name the ion present.

Give the formula for the precipitate.

No precipitate appears when two drops of dilute NaOH solution is added to a colourless solution. More NaOH is added and when heated the gas emitted does not change the colour of red litmus paper.

What is the ion present?

When excess NaOH solution is added to a white precipitate formed by two drops of NaOH, the precipitate remains.

What are the possible ions present?

What are the formulae for these ions?

A white precipitate of a metal is a hydroxide. The precipitate disappears when excess NaOH added. It also disappears when excess ammonia is added.

Name the ion present.

Give the formula for the precipitate.

A white precipitate of this ion does not redissolve in excess NaOH. When dilute sulfuric acid is added to the original solution, a white precipitate forms.

Name the ion.

Give the formulae for the first AND the second precipitates.

A white precipitate of this ion does not redissolve in excess NaOH. When dilute sulfuric acid is added to the original solution, the solution remains colourless.

Name the ion present.

Give the formula for the precipitate.

A white precipitate of this ion redissolves in excess NaOH. When excess dilute ammonia solution is added to the original solution, a white precipitate forms and remains. When dilute sulfuric acid is added to the original solution no precipitate is seen.

Name the ion present.

Give the formula for the first and second precipitate.

A white precipitate of this ion redissolves in excess NaOH. When excess dilute ammonia solution is added to the original solution, a white precipitate forms and remains. When dilute sulfuric acid is added to the original solution, a white precipitate forms.

Name the ion.

Give the formulae for the first, second AND the third precipitates.

A solution turns red litmus paper blue. When dilute HCl is added bubbles of gas are observed.

Name the ion.

Name the gas.

A solution turns red litmus paper blue. When dilute HCl is added no precipitate or gas is seen.

Name the ion.

A solution doesn't change the colour of red litmus paper. When barium nitrate is added a white precipitate forms.

Name the ion.

Give the formula for the precipitate.

A solution doesn't change the colour of red litmus paper. When barium nitrate is added no precipitate forms. When silver nitrate is added to a new solution, no precipitate is formed.

Name the ion.

A solution doesn't change the colour of red litmus paper. When barium nitrate is added no precipitate forms. When silver nitrate is added to a new solution, a precipitate is formed. This precipitate disappears when excess ammonia is added.

Name the ion.

Give the formula for the precipitate.

A solution doesn't change the colour of red litmus paper. When barium nitrate is added no precipitate forms. When silver nitrate is added to a new solution, a precipitate is formed. This precipitate remains when excess ammonia is added.

Name the ion present.

Give the formula for the precipitate.

What is the formula for the sulfate ion?

What is the formula for the nitrate ion?

What is the formula for the chloride ion?

What is the formula for the hydroxide ion?

What is the formula for the carbonate ion?

What is the formula of the ammonium ion?

A blue precipitate appears when two drops of dilute NaOH solution is added to a blue solution.

Write the ionic equation for the formation of the precipitate.

An orange/brown precipitate appears when two drops of dilute NaOH solution is added to an orange solution.

Write the ionic equation for the formation of the precipitate.

A greenish precipitate appears when two drops of dilute NaOH solution is added to a pale green solution.

Write the ionic equation for the formation of the precipitate.

A brown/grey precipitate appears when two drops of dilute NaOH solution is added to a colourless solution.

Write the ionic equation for the formation of the precipitate.

A white precipitate which is a hydroxide disappears in excess NaOH, and also when excess ammonia is added.

Write the ionic equation for the formation of the precipitate.

When excess NaOH solution is added to a white precipitate formed by two drops of NaOH, the precipitate remains.

Write the ionic equations for the formation of all of the possible precipitates.



A white precipitate of an ion does not redissolve in excess NaOH. When dilute sulfuric acid is added to the original solution, a white precipitate forms.

Write the ionic equations for the formation of the precipitates.

A white precipitate of this ion does not redissolve in excess NaOH. When dilute sulfuric acid is added to the original solution, the solution remains colourless.

Write the ionic equation for the formation of the precipitate.

A white precipitate of this ion redissolves in excess NaOH. When excess dilute ammonia solution is added to the original solution, a white precipitate also forms and remains. When dilute sulfuric acid is added to the original solution no precipitate is seen.

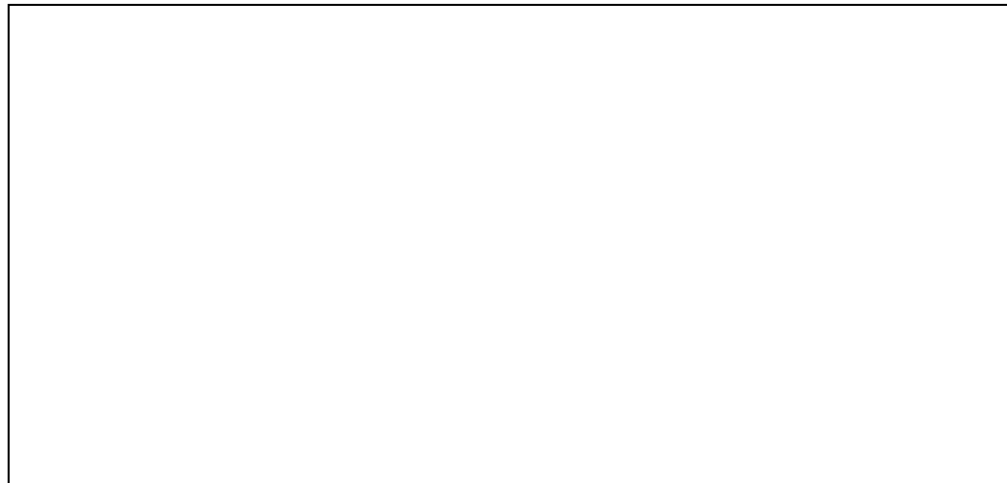
Write the ionic equations for the formation of the precipitates.

A white precipitate of an ion redissolves in excess NaOH. When excess dilute ammonia solution is added to the original solution, a white precipitate forms and remains. When dilute sulfuric acid is added to the original solution, a white precipitate forms.

Write the ionic equations for the formation of the precipitates.

Solution doesn't change the colour of red litmus paper. When barium nitrate solution is added to it a white precipitate forms.

Write the ionic equation for the formation of the precipitate.



A solution doesn't change the colour of red litmus paper. When barium nitrate is added no precipitate forms. When silver nitrate solution is added to the original solution, a white precipitate forms, which dissolves in excess ammonia solution.

Write the ionic equation for the formation of the precipitate.

A solution doesn't change the colour of red litmus paper. When barium nitrate is added no precipitate forms. When silver nitrate solution is added to the original solution, a white precipitate forms, which does not dissolve in excess ammonia solution.

Write the ionic equation for the formation of the precipitate.

A white precipitate which is a hydroxide disappears in excess NaOH, and also when excess ammonia is added to the original solution.

Write the ionic equation for both of the dissolution reactions.

A white precipitate of this ion redissolves in excess NaOH. When excess dilute ammonia solution is added to the original solution, a white precipitate forms and remains. When dilute sulfuric acid is added to the original solution no precipitate is seen.

Write the ionic equations for the dissolution of the precipitate.

A solution doesn't change the colour of red litmus paper. When barium nitrate is added no precipitate forms. When silver nitrate is added to a new solution, a precipitate is formed. This precipitate disappears when excess ammonia is added.

Write the ionic equation for the dissolution of the precipitate.

A blue precipitate formed when two drops of ammonia is added to a blue solution. The precipitate redissolves in excess ammonia solution.

Write the equations for both reactions observed.

No precipitate appears when two drops of dilute NaOH solution is added to a clear solution. More NaOH is added and when heated the gas emitted turns red litmus paper blue.

What is the gas produced?

Write an equation for the production of the gas.

A solution turns red litmus paper blue. When dilute HCl is added bubbles of gas are observed.

Write an equation which shows which gas is produced.

An orange solution produces an orange/brown precipitate when dilute sodium hydroxide solution is added. If a solution of potassium thiocyanate is added to the original solution, a dark red solution is formed.

Name the ion present.

Give the equations for the production of the precipitate and the red compound.

An orange solution produces an orange/brown precipitate when dilute sodium hydroxide solution is added. If a solution of potassium thiocyanate is added to the original solution, a dark red solution is formed.

Name the ion present.

Give the equations for the production of the precipitate and the red compound.

